

MedeA QSPR

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1 Key Benefits of MedeA QSPR

- Determination of geometric descriptors for organic molecules
- Evaluation of thermodynamic properties (of organic compounds) based on Group contributions
- Rapid property calculation enabling screening of compounds and/or providing input conditions for molecular simulation Computational Characteristics
- Provides thermophysical properties (based on Joback's Group Contribution method)
- Uses group contributions to compute properties based on correlations

2 Introduction

"MedeA QSPR: Property Prediction using Group Contributions" employs correlations to predict thermophysical properties based on the chemical groups that constitute an organic molecule. The number and type of groups and their contribution to each property, have been developed and described by Joback and Reid [1]. Only structural information of a chemical compound (a molecule) is used to calculate thermophysical and transport properties, by adding group parameters. This method uses a single group list for all eleven properties therefore allowing the calculation of all properties from a single analysis.

3 Molecular Descriptors

There are 42 molecular descriptors (i.e. groups) in the Joback & Reid method [1]. These molecular descriptors are listed below, along with a short description and the name of the "MedeA variable" that is associated with each descriptor. MedeA variables are useful when printing these quantities in user-defined tables, as part of a flowchart involving a QSPR stage.

[1] Kevin G Joback and Robert C Reid, "Estimation of Pure-Component Properties From Group-Contributions," *Chemical Engineering Communications* 57, no. 1 (February 5, 1987): 233-243.

Table1: Descriptors, *MedeA* variables and a short summary of each descriptor.

Descriptor	<i>MedeA</i> variable	Description
NC1	qspr_NC1_calc	-CH ₃ non-ring groups
NC2	qspr_NC2_calc	-CH ₂ non-ring groups
NC3	qspr_NC3_calc	-CH non-ring groups
NC4	qspr_NC4_calc	-C non-ring groups
NC5	qspr_NC5_calc	=CH ₂ non-ring groups
NC6	qspr_NC6_calc	=CH non-ring groups
NC7	qspr_NC7_calc	=C non-ring groups
NC8	qspr_NC8_calc	=C= non-ring groups
NC9	qspr_NC9_calc	#CH non-ring groups
NC10	qspr_NC10_calc	#C non-ring groups
NC11	qspr_NC11_calc	-CH ₂ non-aromatic ring groups
NC12	qspr_NC12_calc	-CH non-aromatic ring groups
NC13	qspr_NC13_calc	-C non-aromatic ring groups
NC14	qspr_NC14_calc	=CH aromatic ring groups
NC15	qspr_NC15_calc	=C aromatic ring groups
NF1	qspr_NF1_calc	F groups
NCI1	qspr_NCI1_calc	Cl groups
NBr1	qspr_NBr1_calc	Br groups
NI1	qspr_NI1_calc	I groups
NO1	qspr_NO1_calc	-O- groups in OH, in alcohols
NO2	qspr_NO2_calc	-O- groups in OH, in phenol
NO3	qspr_NO3_calc	-O- groups in ethers, nonring
NO4	qspr_NO4_calc	-O- groups in ethers, ring
NO5	qspr_NO5_calc	=O groups, nonring
NO6	qspr_NO6_calc	=O groups, ring
NO7	qspr_NO7_calc	=O groups in aldehydes
NO8	qspr_NO8_calc	-O- groups in carboxylic acid groups
NO9	qspr_NO9_calc	-O- groups in carboxylic ester groups
NO10	qspr_NO10_calc	=O groups in other than NO1-NO9 groups
NN1	qspr_NN1_calc	-NH ₂ groups
NN2	qspr_NN2_calc	-NH groups, nonring
NN3	qspr_NN3_calc	-NH groups, ring
NN4	qspr_NN4_calc	-N groups, nonring
NN5	qspr_NN5_calc	-N= groups, nonring
NN6	qspr_NN6_calc	-N= groups, ring
NN7	qspr_NN7_calc	=NH groups
NN8	qspr_NN8_calc	CN- groups
NN9	qspr_NN9_calc	-NO ₂ groups
NS1	qspr_NS1_calc	-SH groups
NS2	qspr_NS2_calc	-S groups, nonring
NS3	qspr_NS3_calc	-S- groups, ring
NH1	qspr_NH1_calc	-H groups

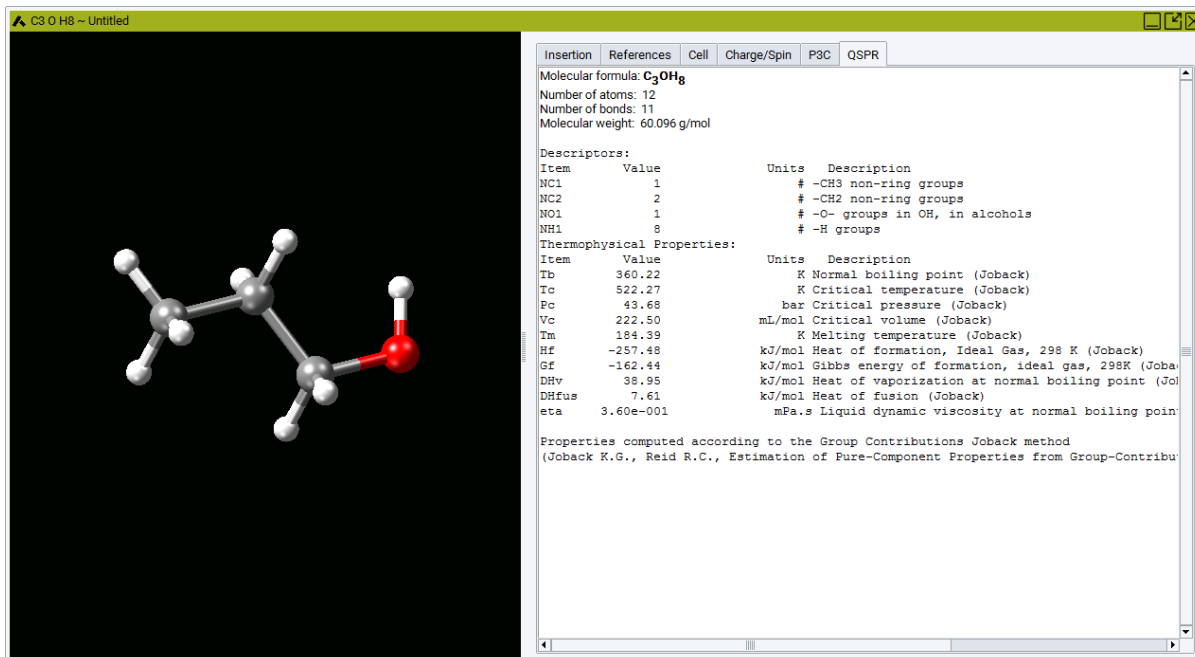
4 Properties

There are ten properties calculated using the Joback & Reid method [1].

Table2: Properties, *Medea* variables and a description of each property.

Property	Units	<i>Medea</i> variable	Description
Tb	K	qspr_Tb_calc	Normal Boiling Point Temperature
Tc	K	qspr_Tc_calc	Critical Temperature
Pc	bar	qspr_Pc_calc	Critical Pressure
Vc	ml/mol	qspr_Vc_calc	Critical Volume
Tm	K	qspr_Tm_calc	Melting Temperature
Hf	kJ/mol	qspr_Hf_calc	Heat of Formation, ideal gas, 298 K
Gf	kJ/mol	qspr_Gf_calc	Gibbs Energy of Formation, ideal gas, 298 K
DHv	kJ/mol	qspr_DHv_calc	Heat of Vaporization at normal boiling point
DHfus	kJ/mol	qspr_DHfus_calc.	Heat of Fusion
Eta	mPas	qspr_eta_calc	Liquid Dynamic Viscosity at normal boiling point

They are available interactively in the Molecular Builder, or through a QSPR: Property Prediction using Group Contributions stage in a *Medea* flowchart.



C3 O H8 -- Untitled

Insertion References Cell Charge/Spin P3C QSPR

Molecular formula: **C₃OH₈**
 Number of atoms: 12
 Number of bonds: 11
 Molecular weight: 60.096 g/mol

Descriptors:

Item	Value	Units	Description
NC1	1	#	-CH3 non-ring groups
NC2	2	#	-CH2 non-ring groups
NO1	1	#	-O- groups in OH, in alcohols
NH1	8	#	-H groups

Thermophysical Properties:

Item	Value	Units	Description
Tb	360.22	K	Normal boiling point (Joback)
Tc	522.27	K	Critical temperature (Joback)
Pc	43.68	bar	Critical pressure (Joback)
Vc	222.50	mL/mol	Critical volume (Joback)
Tm	184.39	K	Melting temperature (Joback)
Hf	-257.48	kJ/mol	Heat of formation, Ideal Gas, 298 K (Joback)
Gf	-162.44	kJ/mol	Gibbs energy of formation, ideal gas, 298K (Joback)
DHv	38.95	kJ/mol	Heat of vaporization at normal boiling point (Joback)
DHfus	7.61	kJ/mol	Heat of fusion (Joback)
eta	3.60e-001	mPa.s	Liquid dynamic viscosity at normal boiling point (Joback)

Properties computed according to the Group Contributions Joback method
 (Joback K.G., Reid R.C., Estimation of Pure-Component Properties from Group-Contribution)

The results for propanol appear like this on the QSPR tab of the Molecular Builder and the output of a QSPR: Property Prediction using Group Contributions stage looks like this in *Job.out*.

Stage 1: Property Prediction using Group Contributions

Molecular formula: C30H8
 Number of atoms: 12
 Number of bonds: 11
 Molecular weight: 60.096 g/mol

Descriptors:

Item	Value	Units	Description
NC1	1	#	-CH3 non-ring groups
NC2	2	#	-CH2 non-ring groups
NO1	1	#	-O- groups in OH, in alcohols
NH1	8	#	-H groups

Thermophysical Properties:

Item	Value	Units	Description
Tb	360.22	K	Normal boiling point (Joback)
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 (Joback K.G., Reid R.C., Estimation of Pure-Component Properties from Group-Contributions, Chem. Eng. Commun. 57, p. 233-243 (1987))

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